



**AMPERE NMR School**

19 – 28 June 2008  
Poznań, Wierzba  
Poland

# ABSTRACTS

Edited by S. Jurga

UNDER THE AUSPICES OF THE AMPERE GROUP  
& ADAM MICKIEWICZ UNIVERSITY, POZNAŃ



**PROGRAMME AMPERE NMR SCHOOL, WIERZBA 2008**

	SUNDAY	MONDAY	TUESDAY	WEDNESDAY	THURSDAY	FRIDAY	SATURDAY		
8.00-9.00	ARRIVAL AND REGISTRATION	BREAKFAST	BREAKFAST	BREAKFAST	BREAKFAST	BREAKFAST	BREAKFAST		
9.00-9.45		<b>R. Kimmich</b> From the basic equation of motion of molecules to NMR measurands: The harmonic radial potential theory of polymers	<b>J. Fraissard</b> NMR of Physisorbed <sup>129</sup> Xe Used as a Probe to Investigate Porous Solids	<b>Sh. Vega</b> New Aspects of Proton Decoupling in Solid State NMR	<b>F. Fujara</b> Spacially resolved NMR in heavy ion irradiated ionic crystals	<b>C.A. de Lange</b> Scope and limitations of accurate structure determination using liquid-crystal NMR	DEPARTURE		
9.45-10.30		<b>R. Wasylishen</b> Probing nuclear spin-spin coupling tensors in solids	<b>D. Michel</b> NMR on ferroelectric materials with very small sizes and on particles confined in nanoporous matrices	<b>A. Wong</b> Application of solid-state NMR spectroscopy to low gamma quadrupolar nuclei	<b>E. Rössler</b> Molecular Dynamics in Soft and hard confinement – a playground for <sup>31</sup> P NMR	<b>E. Burnell</b> What NMR of solutes in liquid-crystalline solvents can tell about the ordering potential			
10.30-11.00		COFFEE BREAK	COFFEE BREAK	COFFEE BREAK	COFFEE BREAK	COFFEE BREAK			
11.00-11.45		Parallel sessions		Parallel sessions		<b>M. Ernst</b> Spin Diffusion in MAS Solid-State NMR		<b>K. Müller</b> Order and dynamics in disordered solids as evaluated by solid state NMR spectroscopy	<b>J. Stepišnik</b> Constrained molecular self-diffusion in the bulk water measured by NMR
		<b>J. Blicharski</b> Rotational Magnetic Resonance and possibilities of a detection	<b>S. Stapf</b> Spatially resolved monitoring of catalytically activated hydrogen peroxide decomposition – a test case for reaction monitoring by NMR	<b>M. Schönhoff</b> Pulsed Field Gradient NMR studies of molecular exchange in colloidal systems	<b>R. Böhmer</b> Deuteron NMR studies of the dynamics in clathrates				
		<b>D. Kruk</b> Various ways to enhance NMR signals: recent theoretical progress	<b>M. Vogel</b> Mechanisms of Ion Transport in Solid-State Electrolytes: Insights from NMR Multi-Time Correlation Functions	<b>F. Grinberg</b> Diffusion and Structure in Self-assembling Systems Studied by NMR	<b>B. Geil</b> Correlation of primary relaxation and high-frequency modes in supercooled liquids. A Deuteron NMR study	Oral presentations <b>B. Grünberg</b> <b>S. Naumov</b> <b>F. Poli</b>		Oral presentations <b>Y.S. Postolenko</b> <b>K. Jasiński</b> <b>G. Woźniak</b>	Oral presentations <b>M. Grbić</b> <b>M. Simčič</b> <b>W. Węglarz</b>
11.45-12.30									
12.30-15.00		LUNCH	LUNCH	LUNCH	LUNCH	LUNCH		LUNCH	
15.00-17.00		Oral presentations <b>I. Rostykus</b> <b>S. Dekarchuk</b> <b>T. Mykhailova</b> <b>S. De Santis</b>	Oral presentations <b>B. Blicharska</b> <b>J. Tritt-Goc</b> <b>S. Poberezhets</b> <b>L. Lalowicz</b> <b>L. Latanowicz</b>			Poster presentations (1-23)		Poster presentations (1-22)	
17.00-19.30		<b>Workshop</b> <b>M. Giersig</b> Nanomaterials and their Applications in Electronic and Biomedicine	<b>Workshop</b> <b>F. Fujara,</b> <b>D. Kruk,</b> <b>E. Rössler</b> Perfect recipe for dealing with strange relaxation data	<b>Workshop</b> <b>R. Wasylishen,</b> <b>D. Michel</b> NMR of Quadrupolar Nuclei	<b>Workshop</b> <b>E. Burnell,</b> <b>C.A. de Lange</b> NMR of Ordered Liquids	Social event	Poster session I	Poster session II	
19.30-20.30	DINNER	"ALL TOGETHER PARTY"	DINNER	DINNER	DINNER	DINNER			

# ARYLDIKETO ACIDS COMPLEXATION ABILITY AND KETO-ENOL TAUTOMERS RATIO IN PRESENCE OF $Mg^{2+}$ . UV/VIS AND NMR SPECTROSCOPY STUDY IN NONAQUEOUS MEDIA

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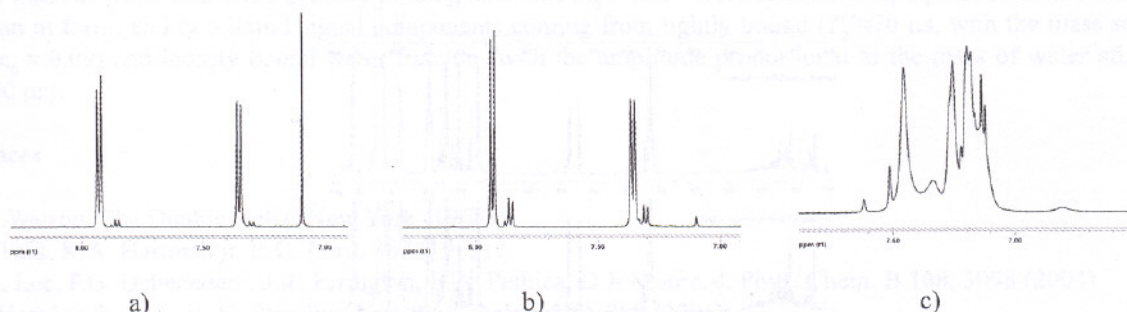
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4-Aryl/heteroaryl-2,4-dioxobutanoic acid (ADK) derivatives exert widespread biological activities. Targeting HIV-1 integrase, the enzyme responsible for integration of viral DNA in host genome, is among the most important ones [1]. ADK- $Mg^{2+}$  complex formation in the active site is postulated as an important factor that determinates degree of enzyme inhibition [2]. Congeneric set of 4-; 3,4- and 2,5-phenyl substituted ADK was synthesized. During routine characterization, mass spectra obtained by liquid chromatography-electrospray ionization (LC-ESI MS) showed presence of  $2M-1$  and  $2(M-1)+Na$  ions for all compounds. It was observed that  $2(M-1)+Na$  are more intensive than  $2M-1$  peaks in spectra of compounds with 3-alkyl substituents, despite substitution in other positions on the phenyl ring. In turn, in MS spectra of all other studied compounds  $2M-1$  peaks are more intensive. This could indicate significantly better complexation ability of 3-alkyl substituted derivatives and might have pharmacological implications.

We have found that there is no significant complexation ability between ADK and  $Mg^{2+}$  in aqueous solutions (pH range 1-8). To check potential differences in complexation ability between 3- and 4- phenyl substituted ADK,  $^1H$  and  $^{13}C$  NMR spectra were recorded in  $CD_3OD$  with and without  $Mg^{2+}$  ion.  $^1H$  NMR spectra of 4-(4-Methylphenyl)-2,4-dioxobutanoic acid and 4-(2,5-Dimethylphenyl)-2,4-dioxobutanoic acid in presence of  $Mg^{2+}$ , are shown on Figure 1 a) and b).



**Fig.1.** Regions of  $^1H$  NMR spectra of 4-(4-Methylphenyl)-2,4-dioxobutanoic acid alone (a), 4-(4-Methylphenyl)-2,4-dioxobutanoic acid (b) and 4-(2,5-Dimethylphenyl)-2,4-dioxobutanoic acid (c) forty minutes after  $Mg^{2+}$  was added to ADK in 1 : 2 (M : L) molar ratio; recorded in  $CD_3OD$ .

As can be seen on selected spectra regions, there is no significant complexing ability of 4-Me- comparing to 2,5-di-Me- substituted derivative. Another interesting observation, considering spectra of 4-Me- derivative, is that keto-enol ratio is significantly changed when  $Mg^{2+}$  ion is present in solution. This was not observed in  $DMSO-d_6$ .

Job's spectrophotometric method [3, 4] was used to confirm the complex stoichiometry (M:L = 1:2) and to compare abilities of 4-; 3,4-; 2,5-phenyl and  $\beta$ -naphthyl substituted derivatives to complex with  $Mg^{2+}$ .

## References

- [1] R. Dayam, J. Deng, N. Neamati, HIV-1 integrase inhibitors: 2003-2004 update, *Med. Res. Rev.* 26(3) (2006) 271-309.
- [2] C. Marchand, A.A. Johnson, E. Semenova, Y. Pommier, Mechanisms and inhibition of HIV integration, *Drug Discovery Today: Disease Mechanisms*, 3(2) (2006) 253-260.
- [3] P. Job, Formation and stability of inorganic complexes in solution, *Ann. Chim. Phys.* 9 (1928) 113-203.
- [4] K. Hyrose, A Practical Guide for the Determination of Binding Constants, *J. Incl. Phenom. Macrocycl. Chem.* 39 (2001) 193-209.

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